



CC13 Groups in the periodic table

Personal learning checklist | 1 of 1

GCSE Edexcel Combined Science | Chemistry Paper 2

Name _____ Class _____ Date _____

Tick one box for each statement. Red: need help. Amber: partly confident.
Green: can do independently.

I can	Red	Amber	Green
Alkali metals			
Describe group 1 physical properties, outer electrons, ion formation and safe storage.	☐	☐	☐
Describe lithium, sodium and potassium water reactions, observations and balanced product equations.	☐	☐	☐
Predict group 1 reactivity trends and explain electron loss using shells, distance and shielding.	☐	☐	☐
Halogen properties			
Describe halogen diatomic molecules, colours, states and trends in melting and boiling points.	☐	☐	☐
Use temperature data to predict halogen states and physical properties of other group members.	☐	☐	☐
Describe metal-halide and hydrogen-halide formation and the damp-litmus test for chlorine.	☐	☐	☐
Halogen reactivity			
Predict halogen displacement reactions and use observations to compare reactivity.	☐	☐	☐
Explain decreasing halogen reactivity using attraction for an incoming electron, shells and shielding.	☐	☐	☐
Explain halogen displacement by electron transfer and identify oxidised and reduced species (Higher).	☐	☐	☐
Noble gases			
Explain noble-gas low reactivity and monatomic particles using full outer shells, including helium.	☐	☐	☐
Link noble-gas uses to low density, non-flammability and low reactivity.	☐	☐	☐
Describe and predict noble-gas boiling-point and density trends from atomic size.	☐	☐	☐

My next revision priority _____