



CC10 Electrolytic processes

Personal learning checklist | 1 of 1

GCSE Edexcel Combined Science | Chemistry Paper 1

Name _____ Class _____ Date _____

Tick one box for each statement. Red: need help. Amber: partly confident.
Green: can do independently.

I can	Red	Amber	Green
The electrolytic cell			
Explain electrolysis, why electrolytes must be molten or dissolved, and why a DC supply is used.	☐	☐	☐
Identify electrode polarities and explain cation and anion movement, distinguishing ions from electrons in wires.	☐	☐	☐
Electrode reactions			
Explain oxidation and reduction by electron transfer and identify their electrolysis electrodes.	☐	☐	☐
Write electrode half-equations, balancing atoms, charges and electrons (Higher).	☐	☐	☐
Predict products from molten binary ionic compounds, including lead bromide and sodium chloride.	☐	☐	☐
Aqueous products			
Predict aqueous cathode and anode products and explain differences from molten electrolytes.	☐	☐	☐
Describe products from copper chloride, concentrated sodium chloride, sodium sulfate and acidified water.	☐	☐	☐
Describe copper sulfate electrolysis with inert electrodes and explain the fading blue colour.	☐	☐	☐
Identify electrolysis gases using appropriate tests and explain why bubbles alone cannot identify them.	☐	☐	☐
Copper electrodes and practical evidence			
Explain copper-electrode mass changes, stable solution colour and purification using an impure anode.	☐	☐	☐
Plan and interpret electrode-mass investigations with controlled current, time and consistent drying.	☐	☐	☐

My next revision priority _____